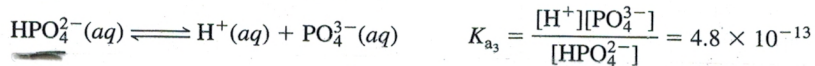
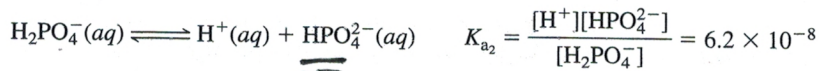
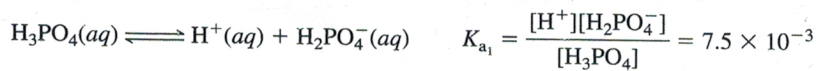


Phosphoric acid (H_3PO_4) is a polyprotic acid with three ionizable hydrogen atoms:



Na_2HPO_4 ... Removing spectator gives us
the anion: HPO_4^{2-} ... I underlined it.

Now simply see if this acts as an
acid or base.

If an acid: this is $K_{a_3} \approx 4.8 \times 10^{-13}$

If a base:



This is reverse of step #2 !! thus get

$$K_{b_2} !! \quad K_{b_2} = \frac{1 \times 10^{-14}}{6.2 \times 10^{-8}} = 1.6 \times 10^{-7}$$

$K_{b_2} > K_{a_3}$... BASIC SALT