Phosphoric acid (H_3PO_4) is a polyprotic acid with three ionizable hydrogen atoms:

$$H_3PO_4(aq) \rightleftharpoons H^+(aq) + H_2PO_4^-(aq) \qquad K_{a_1} = \frac{[H^+][H_2PO_4^-]}{[H_3PO_4]} = 7.5 \times 10^{-3}$$

$$H_2PO_4^-(aq) \rightleftharpoons H^+(aq) + HPO_4^{2-}(aq) \qquad K_{a_2} = \frac{[H^+][HPO_4^{2-}]}{[H_2PO_4^-]} = 6.2 \times 10^{-8}$$

$$HPO_4^{2-}(aq) \rightleftharpoons H^+(aq) + PO_4^{3-}(aq) \qquad K_{a_3} = \frac{[H^+][PO_4^{3-}]}{[HPO_4^{2-}]} = 4.8 \times 10^{-13}$$

Na_2HPO_4 .... Removing spectator gives us the ANION : HPO_4^{--} ... I underlined it.

Now simply see if this acts as AN Acid or base.

If AN Acid : this is $K_{a_3} \approx 4.8 \times 10^{-13}$

If a base :

$$HPO_4^{--} + H^+ \longrightarrow H_2PO_4^-$$

This is reverse of step #2 !! thus Get K_{b_2} !! $K_{b_2} = \frac{1 \times 10^{-14}}{6.2 \times 10^{-8}} = 1.6 \times 10^{-7}$

$K_{b_2} > K_{a_3}$ ... BASIC SALT.